

- xv Screening or shielding effect depends upon no. of _____.
- a) Inner shells b) No. of valence electrons c) None
- xvi Shielding or screening effect _____ in period.
- a) Increase b) Decrease c) Remains Constant
- xvii Shielding effect increases in group, because of increase in _____.
- a) Atomic mass b) Atomic No. c) Inner shell
- xviii What is the increasing order of shielding effect in K, O and Li?
- a) $O > K > Li$ b) $K > Li = O$ c) None
- xix Atomic radii or atomic size of an atom depends upon no. of _____.
- a) No. of shells b) Shielding effects c) Both
- xx Atomic size increases in group because of increase in _____.
- a) No. shells b) Shielding effects c) Both
- xxi. Ionization energy explain the formation of _____.
- a) Cation b) Anion c) Neutral atom
- xxii . Ionization energy is measure in _____.
- a) Calories b) Kilo Joules c) Kilo joule per mole
- xxiii . Ionization energy _____ from bottom to top in a group.
- a) Increase b) Decrease c) Constant
- xxiv . Value of ionization energy determines the _____ of bond.
- a) Polarity b) Strength c) Type
- xxv. Ionization energy is an _____ process.
- a) Exothermic b) Endothermic c) None
- xxvi. High value of ionization energy shows that the bond is _____.
- a) Weak b) Strong c) Non-polar
- xxvii. Ionization energy _____ across a period.
- a) Increase b) Decrease c) Constant
- xviii. Ionization energy value is influence by the _____.
- a) Atomic size b) Shielding effect c) Both
- xxix. The Element with highest ionization energy value is _____.
- a) H b) He c) Fr
- xxx. Electron affinity explains the _____ formation.
- a) Cation b) Anion c) None
- xxxi. Electron affinity is inverse of _____.
- a) Shielding effect b) Ionization energy c) Atomic size
- xxxii. Electron affinity is measure in _____.
- a) Kj/mole b) Calorie c) Kg/mole
- xxxiii. Electron affinity is an _____ process.
- a) Endothermic b) Exothermic c) Both
- xxxiv. Value of electron affinity is always written with _____.
- a) Positive b) Negative c) None
- xxxv. Electron affinity _____ in a group.

- a) Decrease b) increase c) Both
- xxxvi. Electron affinity is influence by _____.
- a) Atomic size b) Shielding effect c) Both
- xxxvii. Electronegativity _____ across a period.
- a) Decrease b) increase c) None
- xxxviii. Electronegativity _____ across a group.
- a) Increase b) Decrease c) None
- xxxix. Electronegativity influence by _____.
- a) Ionization energy b) Electron affinity c) Atomic size
- xl. Electronegativity scale was develop by _____.
- a) Arrhineus b) Rutherford c) Pauling
- xli. Pauling scale value is _____.
- a) 0-4 b) 1-4 c) 0-3
- xl.ii. The most electronegative atom is _____.
- a) Cl b) F c) H
- xl.iii. The electronegativity difference between bonded atom cause _____ in a bond.
- a) Polarity b) Non-polarity c) None
- xl. iv. Reactivity of a metals depend upon _____.
- a) Ability to lose electron b) Valence Shell c) Ability to gain electron
- xl. v. Reactivity of a non-metals depend upon _____.
- a) Ability to lose electron b) Ionization energy c) Ability to gain electron
- xl. vi. Reactivity of metals _____ in a group.
- a) Increase b) Decrease c) None
- xl. vii. Reactivity of metals _____ across a period.
- a) Increase b) decrease c) Both
- xl. viii. Reactivity of metals decrease across a period due _____.
- a) Increase in atomic size b) Decrease in atomic size c) None
- xl. ix. Reactivity of metal decrease in groups due to _____.
- a) Increase in atomic size b) Decrease in atomic size c) None
- l. Reactivity of non-metal depends upon ability to _____ electron easily.
- a) Gain b) Lose c) None
- li. Reactivity of non-metals _____ in a group.
- a) Increase b) Decrease c) None
- lii. Reactivity of non-metals _____ across a period.
- a) Increase b) Decrease c) Both
- liii. Reactivity of non-metals depends upon _____.
- a) Atomic size b) Electronegativity c) Both
- liv. Metallic character define as ability of an atom to _____ electron easily.
- a) Lose b) Gain c) Share
- lv. Metallic character depends upon _____.

- a) Electronegativity b) Electron affinity c) Ionization energy
- lvi. Metallic characters increase in _____.
- a) Group b) Period c) Block
- lvii. Metallic character decrease in _____.
- a) Group b) Period c) Both
- lviii. Metallic character increase in group due to _____.
- a) Increase of atomic size b) Decrease in atomic size c) None
- lix. Metallic character decrease across a period due to _____.
- a) Increase in atomic size b) Decrease in atomic size c) None
- lx. Metallic character is associated with _____.
- a) Metals b) Non-metals c) Metalloids
- lxi. Elements in a similar group have same _____.
- a) Valence shell electronic configuration
b) Electronic configuration
c) Both
- lxii. Elements in group 3 to 12 are called _____.
- a) Normal Elements b) Transition elements c) None
- lxiii. Transition elements have _____ oxidation states.
- a) Similar b) Variable c) Both
- lxiv. _____ have high density due to high atomic mass and compact structure.
- a) Normal elements b) Transition Elements c) Halogen group
- lxv. Transition elements have _____ melting points.
- a) Low b) high c) Moderate
- lxvi. Group 17 (VIIA) is called _____.
- a) Alkali metal b) Noble gas c) Halogen
- lxvii. Halogen group is called halogen because they form salt, when react with _____.
- a) Alkali metal b) Alkaline earth metal c) Nobel gases
- lxviii. Halogen group is also called _____.
- a) Inert group b) Reactive Group c) Salt former
- lxix. In halogen group _____ is radioactive.
- a) Astatin (At) b) Tennessine (Ts) c) Both
- lxx. All halogens have valence shell electronic configuration _____.
- a) $ns^2 np^4$ b) $ns^2 np^5$ c) $ns^2 np^6$
- lxxi. Halogens are highly _____.
- a) Reactive b) Inert c) None
- lxxii. Reactivity of halogen related to their ability of _____ electron.
- a) Gain b) Lose c) Share
- lxxiii. Halogens have strong _____ power.
- a) Oxidizing b) Reducing c) Both
- lxxiv. Which halogen has strongest oxidizing power?
- a) F b) Br c) I

