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a) 8 b) 18 c) 2 ii The second row of the periodic table is also called a) Short b) Long c) Very short iii Elements of group VIII A are called a) Noble gas b) Inert gas c) Both iv Atomic number of an element represents the number of in an atom. a) Protons b) Electrons c) Both v According to Aufbau's principle, which orbital fill first? a) 3s b) 4s c) 2s vi Period of an elemnt is determine by a) Azimithal Quantum no. b) Principal Quantum no. c) Spin Quantum no vii Valence shell electronic configuration of K ₁₉ is a) 4s ¹ b) 3s ² 3p ⁶ c) 3s ² 3p ⁶ , 4s ¹					
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a) 4s ¹ b) 3s ² 3p ⁶ c) 3s ² 3p ⁶ , 4s ¹					
viii Floments of group 12 to 19 org p block alements except					
viii Elements of group 13 to 18 are p-block elements except					
a) C b) He c) Ne					
ix Elements of group IA are called alkali metals because they form, in					
water.					
a) Base b) Acid c) Salt					
x Elements of group IIA are called alkaline earth metals because theirare					
basic in nature and mostly found in					
a) Oxides, Ocean b) Oxides, earth crust c) Oxides, earth crust					
xi Elements of group VIIA are called halogen or salt former, because they form salt					
with elements of					
a) IA ECHEM b) IIA DE TRANSCO IIIA					
xii Elements of group VIIIA are called nobel gas or inert gas because they					
a) They react by itself b) they never react by itself c) None					
xiii Chemical properties of an element depends upon					
a) Electronic configuration b) Valence Shell Electronic Configuration c)					
None					
xiv Physical properties of an element depends upon a) Valence electrons b) Atomic radii c) No of shells					

xv Screening or shielding effe	ct depends upon no. of	•
a) Inner shells	b) No.of valence electrons	c) None
xvi Shielding or screening effe		
a) Increase	b) Decrease	c) Remains Constant
xvii Shielding effect increases i	n group, because of increase	e in
a) Atomic mass	b) Atomic No.	c) Inner shell
xviii What is the increasing orde	er of shielding effect in K, O a	nd Li?
a) O>K>Li	b) K>Li=O	c) None
xix Atomic radii or atomic size	of an atom depends upon no	o. of
a) No. of shells	b) Shielding <mark>effects</mark>	c) Both
xx Atomic size increases in gr	oup because of increase in_	
a) No. shells	b) Shielding effects	c) Both
xxi. Ionization energy explain th	ne formation of	
a) Cation	b) Anion	c) Neutral atom
xxii .lonization energy is measu	ıre in	
a) Calories	b) Kilo Joules	c) Kilo joule per mole
xxiii .lonization energy	from bottom to top in a g	roup.
a) Increase	b) Decrease	c) Constant
xxiv . Value of ionization energy	determines the	_ of bond.
a) Polarity	b) Strength	c) Type
xxv. Ionization energy is an	process.	
a) Exothemic	b) Endothermic	c) None
xxvi. High value of ionization en	ergy shows that the bond is _	
a) Weak	c) Strong	c) Non-polar
xxvii. Ionization energy	across a period.	
a) Increase	c) Decrease	c) Constant
xviii. Ionization energy value is ir	nfluence by the	
a) Atomic size	b) Shielding effect	c) Both
xxix. The Element with highest ic	onization energy value is	
a) H	b) He	c) Fr
xxx. Electron affinity explains th	neformation.	
a) Cation	b) Anion	c) None
xxxi. Electron affinity is inverse o	of	
a) Shielding effect	b) Ionization energy	c) Atomic size
xxxii. Electron affinity is measur	e in	
a) Kj/mole	b) Calorie	c) Kg/mole
xxxiii. Electron affinity is an	process.	
a) Endothermic	c) Exothermic	c) Both
xxxiv. Value of electron affinity is	always written with	•
a) Positive	c) Negative	c) None
xxxv. Electron affinity	in a group.	

	a)	Decrease	b) increase	c) Both
xxxvi. Electron affinity is influence by				
	a)	Atomic size	b) Shielding effect	c) Both
xxxvii.	Electr	onegativity	_ across a period.	
	a)	Decrease	b) increase	c) None
xxxviii.	Elect	ronegativity	across a group.	
	a)	Increase	b) Decrease	c) None
xxxix.	. Electr	onegativity influence	by	
	a)	Ionization energy	b) Electron affinity	c) Atomic size
xl. E	lectror	negativity scale was d	levelop by	
	1	Arrhineus		c) Pauling
xli.	Paulin	g scale value is		
	a)	0-4	b) 1-4	c) 0-3
xlii.	The m	ost electronegative a	tom is	
		Cl	b) F	c) H
xliii.		ectron <mark>egativi</mark> ty differe	ence between bonded atom	causein a
	bond.			
	•		b) Non-polarity	c) None
xliv.			nd upon	
			on b) Valence Shell	
xlv.			depend upon	
			on b) Ionization energy	c) Ability to gain electron
xlvi.		vity of metals		67
		Increase	b) Decrease	c) None
xlvii.		vity of metals	across a period.	
		Increase	b) decraese	c) Both
xlviii.			se across a period due	
			size b) Decrease in atom	
XlIX.			e in groups due to	
	•		size b) Decraese in atom	
ι.			ends upon ability to	
	•	Gain	b) Lose	c) None
U. F	reactivi	ity of non-metals	b) Decrease	COM
I:: F				c) None
ui. F			across a period.	a) D a + b
1:::	,	Incraese	•	c) Both
liii.		-	epends upon	a) Dath
lisz N	•		b) Electronegativity	•
uv. N			ability of an atom to	
ls e	,	Lose	b) Gain	c) Share
lv.	metall	ic character depends	upon	

	а) Electronegativity	b) Electron affinity	c) lonization energy
lvi.	. Metallic	characters increase i	n	
	a)	Group	b) Period	c) Block
lvii	i. Metallic	character decrease i	n	
	a)	Group	b) Period	c) Both
lviii.	Metal	llic character increase	e in group due to	
	a)	Increase of atomic s	size b) Decrease in atom	ic size c) None
lix.	Metallic	character decrease a	cross a period due to	·
	a)	Increase in atomic s	ize b) Decrease in atom	ic size c) None
lx.	Metallic	character is assicitae	ed with	
	a)	Metals	b) Non-metals	c) Metalloids
lxi.	Element	s in a similar group ha	ave same	
	a)	Valence shell electr	onic configurat	
	b)	Electronic configura	tion	
	c)	Both		
lxii.		s in group 3 to 12 are o		
			b) Transition elements	c) None
lxiii.	Transitio	n elements have	oxidation states.	
	a)	Similar	b) Variable	c) Both
lxiv.			ty due to high atomi <mark>c mass a</mark>	and compact structure.
			b) Transition Elements	c) Halogen group
lxv. 7	Transition	elements have		
		Low	b) high	c) Moderate
lxvi.		(VIIA) is called	- Chaff	
		Alkali metal	b) Noble gas	c) Halogen
			because they form salt, whe	
			b) Alkaline earth metal	c) Nobel gases
lxviii. Ha		oup is also called	hline 789	
	,	Inert group	b) Reactive Group	c) Salt former
lxix. In			radioactve.	
	-	Astatin (At)	b) Tenessine (Ts)	c) Both
lxx. All	_		ectronic configuration	
		ns ² np ⁴	b) ns ² np ⁵	c) ns ² np ⁶
lxxi. Ha	logens ar	_	12 I OINLINE / 03	.COM
	,	Reactive	b) Inert	c) None
lxxII. Re	-	_	eir ability ofel	
1	,	Gain	b) Lose	c) Share
ıxxııı. Ha		ave strong		-) D - 41-
	•	Oxidizing	b) Reducing	c) Both
txxiv. Wh	_	gen has strongest oxid		N I
	a)	F	b) Br	c) I

lxxv.	Free halogen displace the halogen next to it, due to			
	a)	Strong oxiding power	b) Strong ionizing power	c) Both
lxxvi.	Which on	e can't oxidize KBr?		
	a)	F	b) Cl	c) I
lxxvii.	Which on	e of the halogen halid ha	is least thermal stablity?	
	a)	HCl	b) HBr	c) HI
lxxviii.	The correct	ct thermal stability order	r of halogen halid is	·
	a)	HCl>HF>HBr>HI	b) HI>HBr>HCl>HF	c) HF>HCl> HBr>HI
lxxix.	-	is called nobel gases	or inert gases.	
	a)	Group 1 (IA)	b) Group 17 (VII A)	c) Group 18 (VIIIA)
lxxx.	Inert gase	s have valence shell ele	ctronic configuration	
	a)	ns ² np ⁶	b) ns ² np ⁵	c) ns ¹ np ⁶
lxxxi.	Which no	bel gas has exceptional	valence shell electronic Con	figuration?
	a)	He	b) Ne	c) Ar
xxxii. I	Metals are	d <mark>uctile and mallea</mark> ble, d	lue to	
	a)	lonic bond	b) Covalent bond	c) metallic bond
xxxiii.	Which no	n-metal is good conduct	tor of electricity?	
	a)	Graphite	b) Sulphur	c) Phosphorrous
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